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Quantities Name _____

10.1 The Mole: A

Measurement of Matter

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1. What is the mole? 2. What is Avogadro's number? 3. What kinds of things can you count using the mole? 4. What are three ways of measuring the amount of a substance? 5.

Guided Notes

Chapter 10:

Chemical Quantities

Name 10.1 The ...

CHAPTER 10: Chemical
Quantities BASICS: •

The basic unit that is
used to determine the

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amount of a chemical substance is called a mole • A mole (mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

Chapter 10 Chemical
Quantities 91. SECTION
10.1 THE MOLE: A

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CHEMICAL QUANTITIES
MEASUREMENT OF MATTER (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

Measuring Matter
(pages 287-289)

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages

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287-296)

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Answers to Mole
Conversions

Worksheet. Chapter 7

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Section 7.1 The Mole:

A. rough. File.

Introduction to SOL

Review. Stoichiometry -

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Masses Worksheet.

unit 2b pp4.

chapter 16.1 - Colorado

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Guided Reading ...

Section 10.3 Percent

Composition and

Chemical Formulas p.

325-333 . By the end of

this section you should

be able to: calculate

the percent

composition of a

compound; ...

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chemwins**

Chapter 10 - Chemical
Quantities - 10.1 The
Mole: A Measurement
of Matter - 10.1 Lesson
Check - Page 315: 9
Answer Knowing how
the count mass and
volume of an item
relate a common unit
allows you convert
among these units.

**Chemistry (12th
Edition) Chapter 10 -**
Page 10/26

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Chemical Quantities

...

Section 10.3 – Percent
Composition and

Chemical Formulas.

The percent by mass
(percent composition)
of an element in a
compound is the
number of grams of the
element divided by the
mass in grams of the
compound multiplied
by 100%. % mass of
element = mass of
element x 100. mass of
compound.

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Chapter 9- Chemical Names/Formulas

Basics: Notes, Review

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PHYSICAL QUANTITIES;

The little mole who
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Mole!!) Flashcards
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CourseNotes

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Quantities 247 Practice

Problems In your

notebook, solve the

following problems

SECTION 101 THE

MOLE: A

MEASUREMENT OF

MATTER 1. What is the

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molar mass of sucrose
(C₁₂H₂₂O₁₁)? 2

What is the molar mass
of each of the following
compounds? a
phosphorus
pentachloride (PCl₅) b
uranium hexafluoride
(UF₆)

[Book] Chapter 10 Chemical Quantities 247

this Chapter 10
Chemical Quantities
Packet Answers can be
taken as skillfully as

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quantities
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Here is an example.
Pure HCl (hydrogen chloride) is a gas that is very soluble in water. A plot of the partial pressure of gaseous HCl in equilibrium with aqueous HCl, as a function of the solution molality (Fig. 10.1), shows that the limiting slope at infinite dilution

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is not finite, but zero.

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