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following questions in the space provided 1 a Write the two equations that show the two-stage ionization of sulfurous acid in water stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \sim \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b Which stage of ionization usually produces more ions?

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1 . One example for  $\text{NH}_3$  as a conjugate acid:  $\text{NH}_3 + \text{H}^+ \rightarrow \text{NH}_4^+$ ; as a conjugate base:  $\text{NH}_4^+ + (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{NH}_3 (\text{aq}) + \text{H}_2\text{O}$ . (a)  $\text{NaHSeO}_3 < \text{NaHSO}_3 < \text{NaHSO}_4$ ; in polyoxy acids, the more electronegative central element—S, in this case—forms the stronger acid. The larger number of oxygen atoms on the central atom (giving it a higher oxidation state) also creates ...

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Download File PDF Chapter 14 Chemistry Answers Section Assessment 14.1 page 479 1. Explain Use the properties of seawater to describe the characteristics of mixtures. Answers will vary but might include that Mixtures and Solutions Answer Questions, Exercises, and Problems 14–15. Check your answers with those at the end of the chapter.

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CHAPTER 14 REVIEW Acids and Bases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Name the following compounds as acids: sulfuric acid a.  $\text{H}_2\text{SO}_4$  sulfurous acid b.  $\text{H}_2\text{SO}_3$  hydrosulfuric acid c.  $\text{H}_2\text{S}$  perchloric acid d.  $\text{HClO}_4$  hydrocyanic acid e. hydrogen cyanide 2.  $\text{H}_2\text{S}$  Which (if any) of the acids mentioned in item 1 are binary acids? 3.

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### **Chapter 14 Chemical Kinetics**

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Chemistry 102 ANSWER KEY REVIEW QUESTIONS Chapter 14 1. A mixture of 0.10 mol of NO, 0.050

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mol of H<sub>2</sub> and 0.10 mol of H<sub>2</sub>O is placed in a 1.0L flask and allowed to reach equilibrium as shown below:  $\text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g}) \rightleftharpoons 2 \text{NO}(\text{g}) + 2 \text{H}_2(\text{g})$  At equilibrium [NO] = 0.062 M. Calculate the equilibrium constant, K<sub>c</sub>, for this reaction.

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